



Part 1 - Compulsory

Section A GCSE Review

Q4.1-4.2

This question is about the halogens.

1 Write the State symbol for chlorine at room temperature. [1 mark]

Cl<sub>2</sub> ( \_\_\_\_\_ )

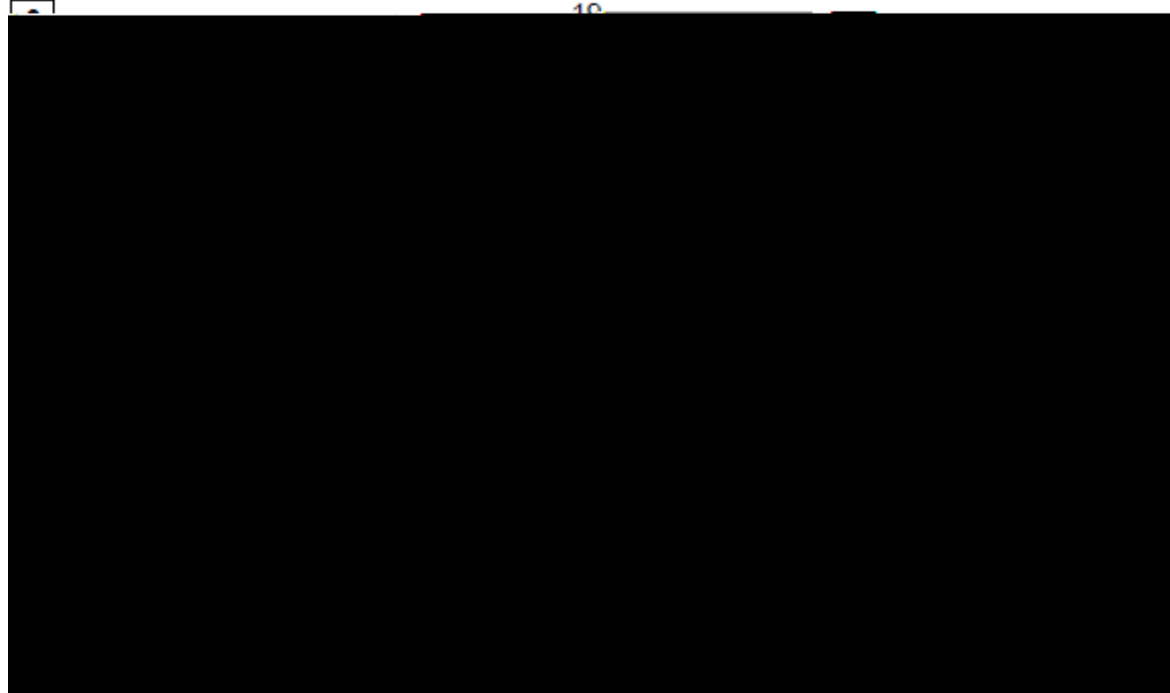
2 Figure 4 represents one molecule of fluorine.

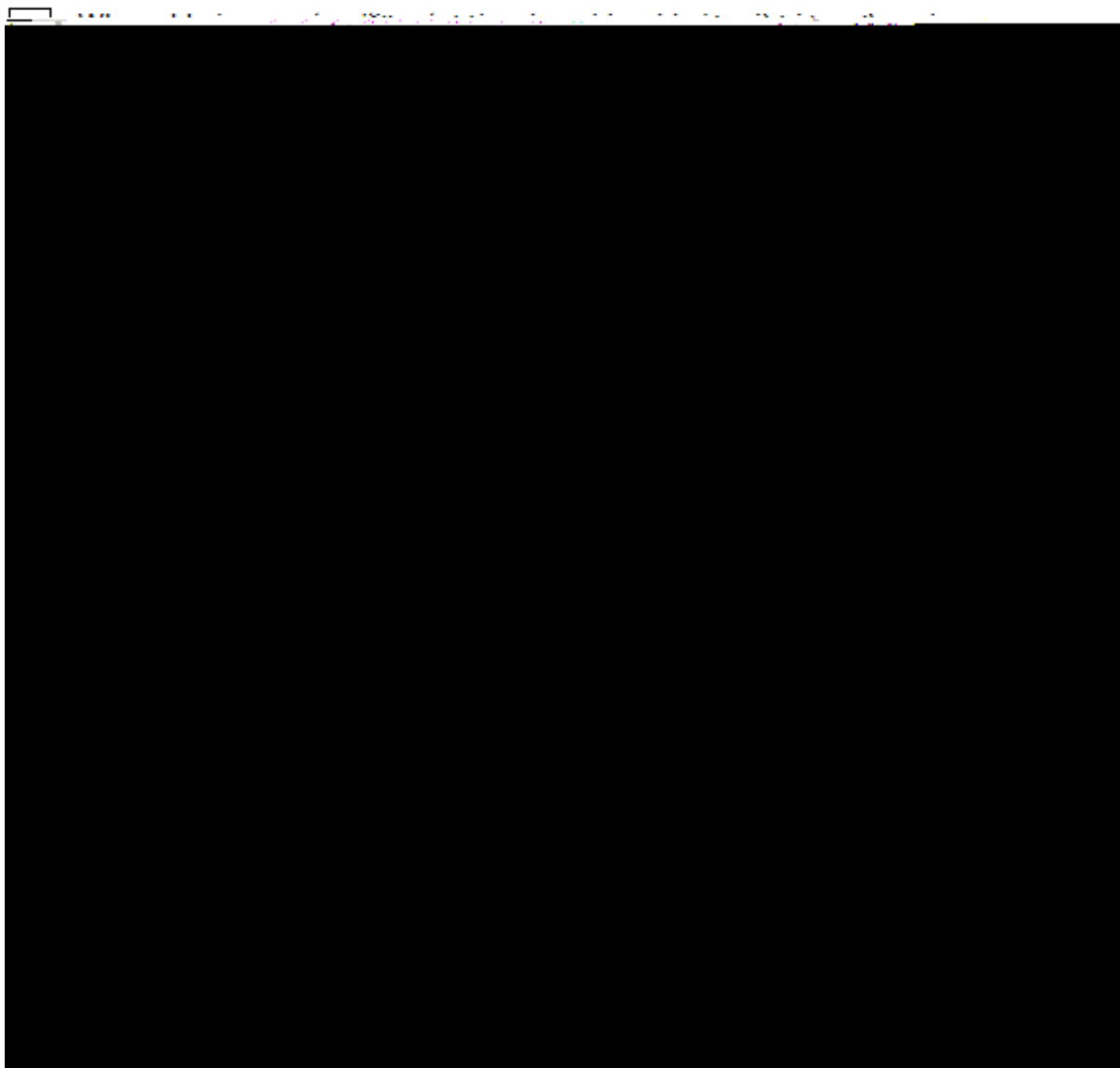
2 Figure 4 represents one molecule of fluorine. [2 marks]

Figure 4 is a dot-and-cross diagram for a molecule of fluorine. It shows two fluorine atoms, each with a full outer shell of eight electrons. The two atoms are joined together by a single bond, which is represented by a pair of shared electrons (one from each atom). The diagram shows two fluorine atoms, each with a full outer shell of eight electrons. The two atoms are joined together by a single bond, which is represented by a pair of shared electrons (one from each atom).



Q4.3-4.5





Q6.1, 6.2

is heated.

1

Give the formula of each product when calcium carbonate ( $\text{CaCO}_3$ )

12 marks

and

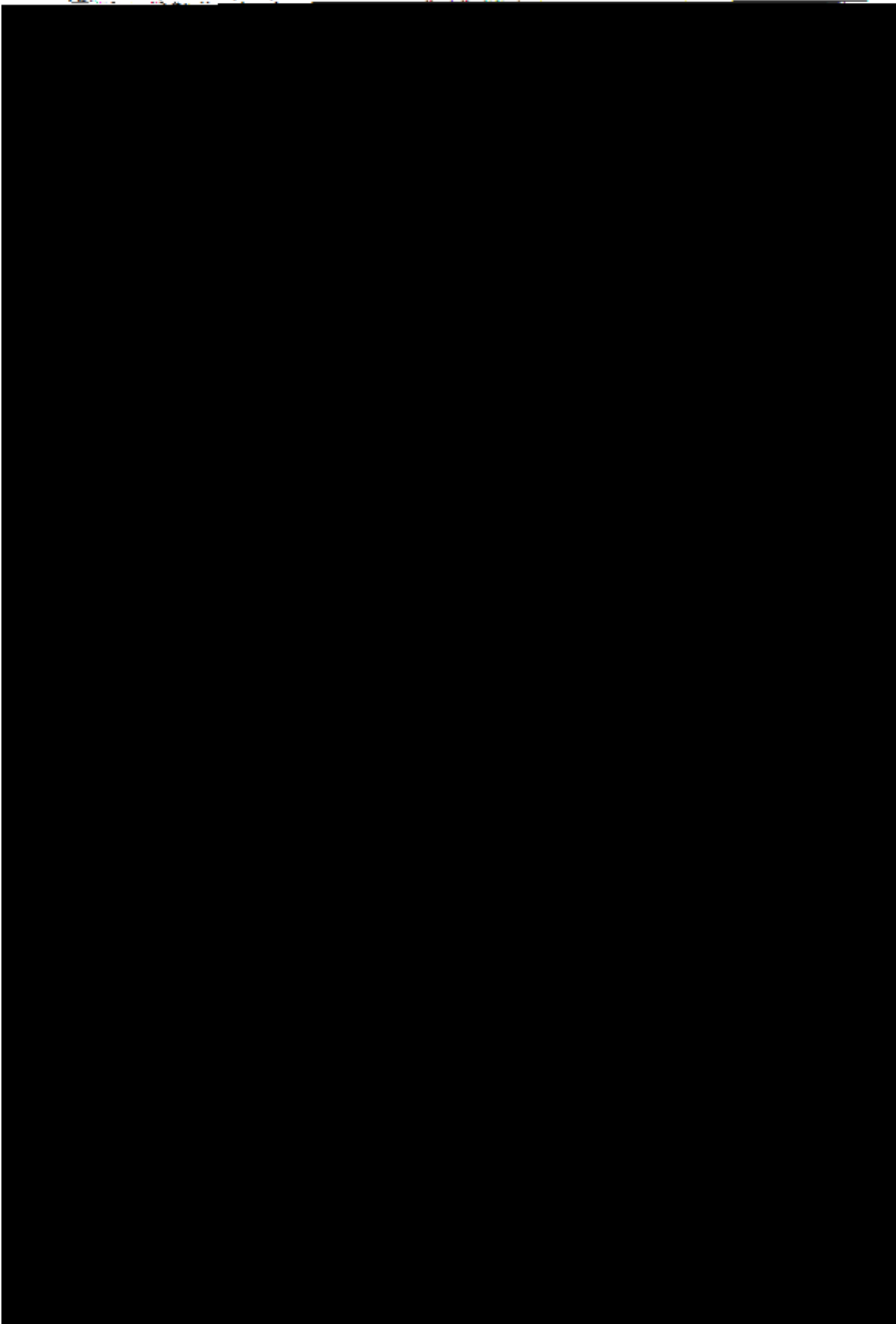
the relative atomic mass ( $A_r$ ) of the Group 2 metal in the metal carbonate. Calculate

Name the Group 2 metal

(2 marks)

Metal

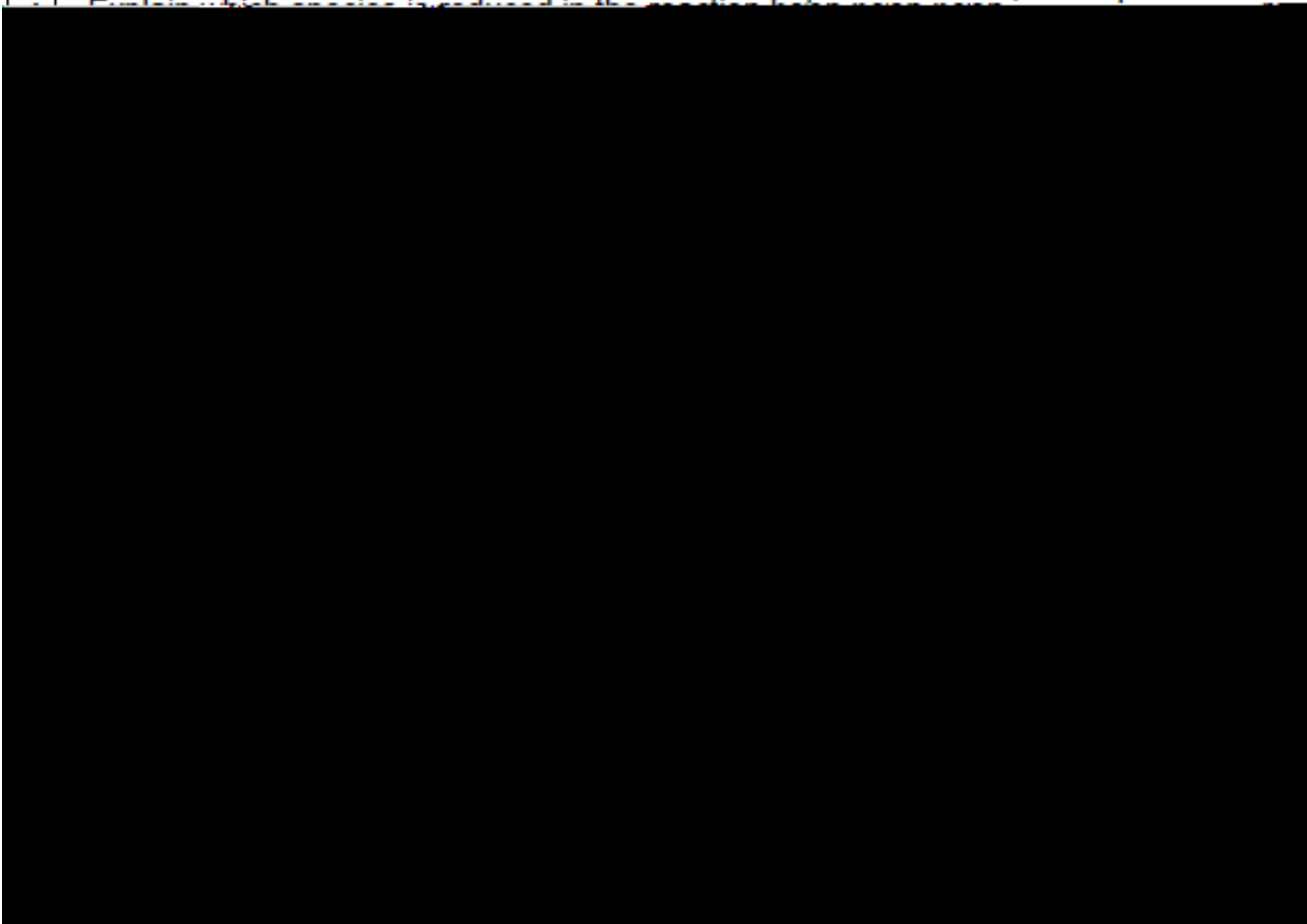
Q8.1-8.3



Q8.4



Can you identify any species introduced in the region by humans?









Task 1



Task 4







Task 8

\_\_\_\_\_

100  
05

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\_\_\_\_\_  
\_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_









# Significant figures

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not

without

with

is

Task 3.

Original number	3 significant figures	2 significant figures	1 significant figure
2.856			
44.503			
18.29			
0.099			
532.41			



## Decimal places

Task 5.

Original number	2 decimal places	1 decimal place	0 decimal places (an Integer)
12.947			
84.3524			
0.765			

## Standard form

a

a

a

a

Task 6.







Task 9.



least half

-



Figure 2.

(b)

(c)

(1)